

High-Pressure Viscosity and Density Behavior of Ternary Mixtures: 1-Methylnaphthalene + *n*-Tridecane + 2,2,4,4,6,8,8-Heptamethylnonane

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The dynamic viscosity η and the density ρ of the ternary system, *n*-tridecane + 1-methylnaphthalene + 2,2,4,4,6,8,8-heptamethylnonane, were measured as a function of temperature from 293.15 to 353.15 K in 10 K increments at pressures up to 100 MPa. A falling body viscometer was used for measuring the dynamic viscosity above 0.1 MPa, while at 0.1 MPa the viscosity was obtained with an Ubbelohde viscometer. The overall uncertainty in the reported data is less than $1 \text{ kg} \cdot \text{m}^{-3}$ for densities and 2% for viscosities, except at 0.1 MPa where the uncertainty is less than 1%. The experimental results correspond to 882 values of viscosity. With reference to the 126 values published previously for the pure compounds and 882 values for the three associated binaries, the system is globally described by 1890 experimental values as a function of pressure, temperature, and composition. The results for the viscosity are discussed in terms of mixing laws and the excess activation energy of viscous flow.

KEY WORDS: 1-methylnaphthalene; 2,2,4,4,6,8,8-heptamethylnonane; density; excess activation energy of flow; high pressure; hydrocarbon; mixture; *n*-tridecane; viscosity.

1. INTRODUCTION

Our laboratories have been working for several years on the properties of petroleum fluids and synthetic mixtures, more or less representative of

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petroleum fluids, as a function of pressure, temperature, and composition. In petroleum engineering, knowledge of the dynamic viscosity, η , as a function of the reservoir conditions (temperature and pressure, as well as the composition of the fluid) is of major importance because several models, covering a wide range of aspects such as reservoir simulations or designing of transport equipments, use this coefficient. While there are abundant viscosity data for mixtures versus temperature T at atmospheric pressure, studies of viscosity versus pressure P are less frequent, particularly for mixtures and those likely to be used to model a real petroleum fluid or a real petroleum cut, in other words, a complex fluid. It is important to stress here that viscosity data for some binary systems are available in which the viscosity as a function of composition, pressure, and temperature is well described, but there are only a few systematic studies concerning ternary mixtures [see, e.g., Ref. 1 for the ternary n -heptane + methylcyclohexane + 1-methylnaphthalene described by 21 compositions up to 100 MPa (378 experimental data); Ref. 2 for the ternary n -pentane + n -octane + n -decane described by 15 compositions up to 25 MPa (358 experimental data); and Ref. 3 for the ternary water + 2-propanol + diacetone alcohol described by 36 compositions up to 100 MPa (648 experimental data)].

There are many ways to model a real petroleum cut. We decided, some years ago, to choose the ternary system 1-methylnaphthalene + n -tridecane + 2,2,4,4,6,8,8-heptamethylnonane, as part of a very simple representation of some petroleum distillation fractions at 510 K. At atmospheric pressure, the boiling temperature is 507.1 K for n -tridecane, 514.7 K for 1-methylnaphthalene, and 513.1 K for 2,2,4,4,6,8,8-heptamethylnonane. The dynamic viscosity and density for the three pure compounds and the three associated binary systems have already been studied [4–6] up to 100 MPa and in the temperature range 293.15 to 353.15 K, involving 126 points for the pure compounds and 882 points for the three binary systems. This paper extends the study to cover measurements of the dynamic viscosity and density of ternary mixtures under the same pressure and temperature conditions as for the pure compounds and binary mixtures, with a sufficient range of compositions to provide a complete coverage of the representative ternary diagram.

2. EXPERIMENTAL

2.1. Apparatus

The dynamic viscosity η was determined with the aid of a falling body viscometer, details of which are provided in Ref. 7, of the type designed by Ducoulombier et al. [8] (using n -decane and toluene as reference fluids).

Values of the density ρ for pressures between 0.1 and 60 MPa were measured with an Anton-Paar DMA60 resonance densimeter combined with an additional 512P cell. Details of the calibration of this type of apparatus, with vacuum and water as reference fluids, have been described in Ref. 9. The density measurements were extrapolated to 100 MPa according to the procedure described in Ref. 7 using a Tait-like relationship for the variation of the density with pressure:

$$\frac{1}{\rho(P, T)} = \frac{1}{\rho_0(T)} + A(T) \ln \left(1 + \frac{P - P_0}{B(T)} \right) \quad (1)$$

where $\rho_0(T)$ is the density at the pressure P_0 and at the temperature T . Here we chose $P_0 = 0.1$ MPa. For each temperature, the extrapolation of the experimental densities to 80 and 100 MPa was performed by adjusting the A and B parameters. The validity of this method is discussed in Ref. 7 which has been tested with values provided by Dymond et al. [10] for pure alkanes and binary mixtures up to 500 MPa. For example, for isooctane at $T = 348.15$ K, using the parameters A and B adjusted only with density values up to 40 MPa, a small difference of $-0.4 \text{ kg} \cdot \text{m}^{-3}$ (i.e., -0.06%) between experimental and calculated values at 105.2 MPa is obtained, which corresponds approximately to our upper limit of viscosity measurements. It should be emphasized here that an error of $1.0 \text{ kg} \cdot \text{m}^{-3}$ in the sample density generates an error of $1/8000$ in the viscosity [7]. For the viscosity measurements, the uncertainty in the temperature was estimated to be ± 0.5 K, and for the density measurements the uncertainty in the temperature was estimated to be ± 0.05 K. The uncertainty in the pressure was estimated to be ± 0.05 MPa for the density measurements (HBM manometer) and ± 0.1 MPa for the viscosity measurements (SEDEME manometer), except at 0.1 MPa. The overall uncertainty in the reported density values is less than $1 \text{ kg} \cdot \text{m}^{-3}$, while the uncertainty in the viscosity is approximately 2%. As discussed previously [1, 7, 11–13], this error is comparable with the error obtained by other authors for similar experimental systems. Comparative curves have been reported for *n*-heptane and methylcyclohexane in Ref. 13, for water and 2-propanol in Ref. 11, and for 2,2,4,4,6,8,8-heptamethylnonane in Ref. 5, which contain plots of our values and of those obtained by other authors. It should be pointed out that at atmospheric pressure the kinematic viscosity η/ρ was determined with a classical capillary viscometer. For this purpose several Ubbelohde tubes, connected to an automatic AVS350 Schott Geräte analyzer, were used. In this case, the uncertainty in the temperature was ± 0.05 K. After multiplying by the density, the uncertainty of the dynamic viscosity is less than 1%.

2.2. Characteristics of the Samples

The substances used are commercially available chemicals with the following purity levels: *n*-tridecane ($C_{13}H_{28}$; Tokyo Kasei; purity, > 99%; molar mass $M = 184.37 \text{ g} \cdot \text{mol}^{-1}$), 1-methylnaphthalene ($C_{11}H_{10}$; Fluka; purity, > 97%; $M = 142.20 \text{ g} \cdot \text{mol}^{-1}$), and 2,2,4,4,6,8,8-heptamethylnonane ($C_{16}H_{34}$; Aldrich; purity, > 98%, $M = 226.44 \text{ g} \cdot \text{mol}^{-1}$). The mixtures were prepared by very careful weighing (with a Mettler balance) at atmospheric pressure and ambient temperature to obtain the molar fractions $x_i = 0.125, 0.250, 0.375, 0.500, 0.625,$ and 0.750 (with $\sum_{i=1}^3 x_i = 1$) corresponding to the 21 points shown in the ternary diagram (Fig. 1). The systems corresponding to the three summits (pure compounds) and the three sides (binaries) have already been studied [4–6]. The pure fluids, not degassed, were stored in hermetically sealed bottles. The samples were in the liquid state within the experimental temperature and pressure domain.

2,2,4,4,6,8,8-Heptamethylnonane

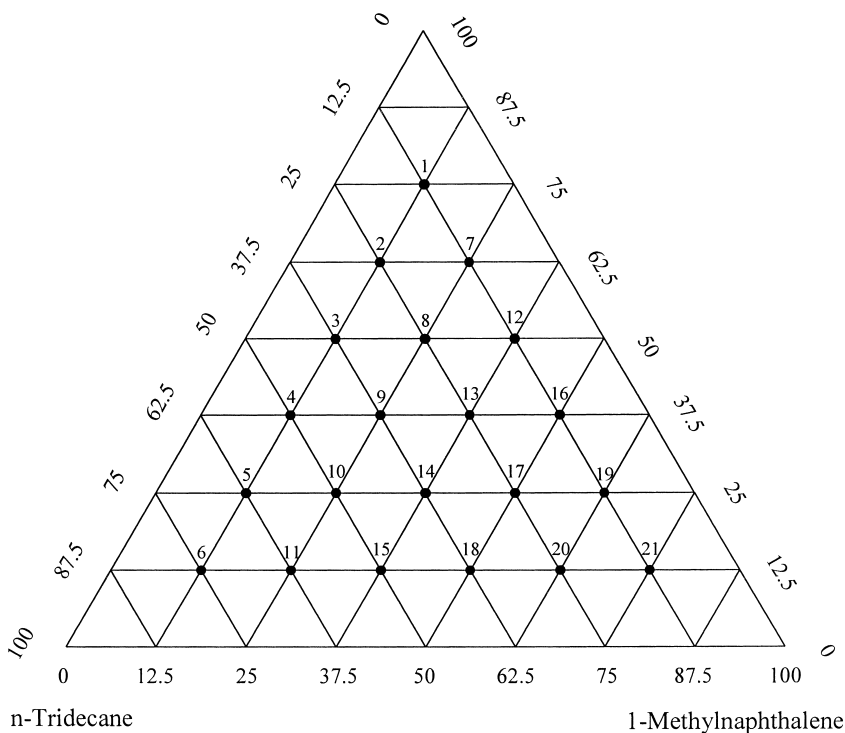


Fig. 1. Ternary mixtures studied in this work represented as points 1–21 in the ternary diagram (versus mol%).

3. RESULTS

Measurements of the dynamic viscosity η were made at seven temperatures (293.15, 303.15, 313.15, 323.15, 333.15, 343.15, and 353.15 K) and at six pressures (0.1, 20, 40, 60, 80, and 100 MPa) for the 21 compositions indicated in Fig. 1. A total of 882 values was obtained for the viscosity. The density measurements were carried out at the same temperatures and compositions at pressures from 0.1 MPa to 60 MPa in steps of 5 MPa (13 pressures), corresponding to 1911 experimental values for the density. These values have been extrapolated (with an error of about 0.1% at $P = 100$ MPa) with the aid [7] of the Tait-type relationship to obtain densities at 80 and 100 MPa (294 values). However, to reduce the amount of tabulated data, we retain only those temperatures and pressures for which the dynamic viscosity has been measured. Additional density data (1029 values) will be published in Refs. 14 and 15. Table I presents the measured dynamic viscosity and density values as a function of temperature T , pressure P , and mole fraction (x_m for 1-methylnaphthalene, x_t for n -tridecane, and x_h for 2,2,4,4,6,8,8-heptamethylnonane; for $x_m, x_t, x_h \neq 0$ and 1). Figures 2 and 3 show the variations of density with pressure (for different temperatures) and with temperature (for different pressures) in the case corresponding to point 9, which is close to the middle of the ternary diagram in Fig. 1 ($x_m = 0.250, x_t = 0.375, x_h = 0.375$). Figures 4 and 5 show the dynamic viscosity under the same conditions. Table I and Figs. 2–5 present a general pattern consistent with previous observations made by other authors and by us on either pure hydrocarbons, binary mixtures, or ternary mixtures of hydrocarbons. The group of isothermal and isobaric curves is regular. This is also true for the density, but in the case of isothermal curves, a concavity is observed associated with a second negative derivative. This form is compatible with the logarithmic form proposed by Tait to model the influence of pressure on $1/\rho$, which is the form used for the extrapolation (see, e.g., Ref. 7). It should be noted that the variations of density with T are practically linear due to the small temperature interval (60 K) in this investigation, because the main aim has been to observe the variations of density and viscosity with pressure and composition.

Figure 6 shows, for $P = 40$ MPa and $T = 323.15$ K, the variations of density with x_t (n -tridecane) for a constant x_m (1-methylnaphthalene). Points M, T, and H correspond to 1-methylnaphthalene, n -tridecane, and 2,2,4,4,6,8,8-heptamethylnonane, respectively, and the sides MT, MH, and TH to the associated binaries. To complete this figure and the following figures, the reported data for the pure compounds and the binaries given in Refs. 4–6 have been used. Within the experimental uncertainty the variations are practically linear, which correspond to very low excess volumes.

Table I. Density ρ and Dynamic Viscosity η Versus Temperature T and Pressure P for Ternary Mixtures of 1-Methylnaphthalene (m) + n -Tridecane (t) + 2,2,4,4,6,8,8-Heptamethylnonane (h)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.125, x_t = 0.125, x_h = 0.750$			
293.15	0.1	0.7964	3.130
293.15	20	0.8089	4.380
293.15	40	0.8197	6.070
293.15	60	0.8292	8.153
293.15	80	0.8377	11.055
293.15	100	0.8454	14.696
303.15	0.1	0.7895	2.504
303.15	20	0.8027	3.481
303.15	40	0.8139	4.661
303.15	60	0.8238	6.221
303.15	80	0.8325	8.258
303.15	100	0.8403	10.896
313.15	0.1	0.7827	2.049
313.15	20	0.7965	2.801
313.15	40	0.8083	3.718
313.15	60	0.8184	4.903
313.15	80	0.8272	6.411
313.15	100	0.8351	8.314
323.15	0.1	0.7759	1.708
323.15	20	0.7905	2.386
323.15	40	0.8027	3.127
323.15	60	0.8133	4.020
323.15	80	0.8226	5.132
323.15	100	0.8310	6.588
333.15	0.1	0.7691	1.447
333.15	20	0.7844	2.002
333.15	40	0.7972	2.585
333.15	60	0.8081	3.287
333.15	80	0.8177	4.127
333.15	100	0.8263	5.179
343.15	0.1	0.7621	1.248
343.15	20	0.7783	1.740
343.15	40	0.7915	2.250
343.15	60	0.8029	2.856
343.15	80	0.8127	3.567
343.15	100	0.8216	4.389
353.15	0.1	0.7552	1.083
353.15	20	0.7721	1.500
353.15	40	0.7860	1.920
353.15	60	0.7977	2.434
353.15	80	0.8079	3.054
353.15	100	0.8169	3.721

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.125, x_t = 0.250, x_h = 0.625$			
293.15	0.1	0.7936	2.775
293.15	20	0.8061	3.894
293.15	40	0.8169	5.210
293.15	60	0.8264	6.900
293.15	80	0.8349	9.042
293.15	100	0.8427	11.735
303.15	0.1	0.7870	2.241
303.15	20	0.8002	3.113
303.15	40	0.8114	4.120
303.15	60	0.8213	5.376
303.15	80	0.8301	6.920
303.15	100	0.8381	8.799
313.15	0.1	0.7802	1.846
313.15	20	0.7940	2.550
313.15	40	0.8057	3.364
313.15	60	0.8159	4.354
313.15	80	0.8249	5.540
313.15	100	0.8330	6.944
323.15	0.1	0.7732	1.550
323.15	20	0.7877	2.128
323.15	40	0.7999	2.781
323.15	60	0.8105	3.568
323.15	80	0.8199	4.501
323.15	100	0.8283	5.594
333.15	0.1	0.7663	1.326
333.15	20	0.7816	1.815
333.15	40	0.7943	2.331
333.15	60	0.8053	2.945
333.15	80	0.8148	3.662
333.15	100	0.8233	4.493
343.15	0.1	0.7593	1.144
343.15	20	0.7755	1.561
343.15	40	0.7887	1.983
343.15	60	0.8000	2.489
343.15	80	0.8099	3.083
343.15	100	0.8187	3.774
353.15	0.1	0.7523	0.999
353.15	20	0.7692	1.375
353.15	40	0.7830	1.740
353.15	60	0.7948	2.167
353.15	80	0.8050	2.658
353.15	100	0.8140	3.216

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.125, x_t = 0.375, x_h = 0.500$			
293.15	0.1	0.7903	2.524
293.15	20	0.8027	3.436
293.15	40	0.8135	4.526
293.15	60	0.8230	5.871
293.15	80	0.8315	7.506
293.15	100	0.8392	9.519
303.15	0.1	0.7833	2.052
303.15	20	0.7965	2.765
303.15	40	0.8076	3.614
303.15	60	0.8176	4.644
303.15	80	0.8263	5.875
303.15	100	0.8343	7.284
313.15	0.1	0.7764	1.702
313.15	20	0.7902	2.306
313.15	40	0.8019	2.990
313.15	60	0.8121	3.801
313.15	80	0.8211	4.749
313.15	100	0.8293	5.819
323.15	0.1	0.7695	1.434
323.15	20	0.7840	1.930
323.15	40	0.7961	2.475
323.15	60	0.8068	3.114
323.15	80	0.8161	3.854
323.15	100	0.8246	4.704
333.15	0.1	0.7625	1.229
333.15	20	0.7778	1.650
333.15	40	0.7905	2.097
333.15	60	0.8014	2.613
333.15	80	0.8110	3.199
333.15	100	0.8196	3.877
343.15	0.1	0.7554	1.066
343.15	20	0.7716	1.414
343.15	40	0.7848	1.783
343.15	60	0.7961	2.213
343.15	80	0.8060	2.706
343.15	100	0.8148	3.279
353.15	0.1	0.7484	0.934
353.15	20	0.7654	1.245
353.15	40	0.7792	1.568
353.15	60	0.7910	1.944
353.15	80	0.8011	2.376
353.15	100	0.8101	2.859

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.125, x_t = 0.500, x_h = 0.375$			
293.15	0.1	0.7872	2.319
293.15	20	0.7997	3.083
293.15	40	0.8105	3.978
293.15	60	0.8201	5.073
293.15	80	0.8286	6.390
293.15	100	0.8363	7.962
303.15	0.1	0.7805	1.889
303.15	20	0.7936	2.551
303.15	40	0.8048	3.259
303.15	60	0.8148	4.092
303.15	80	0.8236	5.056
303.15	100	0.8316	6.158
313.15	0.1	0.7736	1.573
313.15	20	0.7874	2.129
313.15	40	0.7990	2.715
313.15	60	0.8093	3.396
313.15	80	0.8183	4.173
313.15	100	0.8265	5.050
323.15	0.1	0.7666	1.333
323.15	20	0.7811	1.814
323.15	40	0.7932	2.304
323.15	60	0.8038	2.859
323.15	80	0.8131	3.477
323.15	100	0.8215	4.153
333.15	0.1	0.7595	1.149
333.15	20	0.7748	1.544
333.15	40	0.7875	1.936
333.15	60	0.7985	2.387
333.15	80	0.8080	2.896
333.15	100	0.8166	3.465
343.15	0.1	0.7524	0.998
343.15	20	0.7685	1.340
343.15	40	0.7817	1.666
343.15	60	0.7930	2.034
343.15	80	0.8030	2.440
343.15	100	0.8119	2.884
353.15	0.1	0.7453	0.882
353.15	20	0.7622	1.177
353.15	40	0.7759	1.461
353.15	60	0.7877	1.777
353.15	80	0.7979	2.123
353.15	100	0.8070	2.497

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.125, x_t = 0.625, x_h = 0.250$			
293.15	0.1	0.7839	2.141
293.15	20	0.7963	2.841
293.15	40	0.8071	3.606
293.15	60	0.8166	4.518
293.15	80	0.8252	5.587
293.15	100	0.8329	6.829
303.15	0.1	0.7769	1.758
303.15	20	0.7901	2.326
303.15	40	0.8013	2.939
303.15	60	0.8112	3.649
303.15	80	0.8200	4.458
303.15	100	0.8280	5.368
313.15	0.1	0.7700	1.469
313.15	20	0.7838	1.942
313.15	40	0.7954	2.445
313.15	60	0.8057	3.028
313.15	80	0.8147	3.618
313.15	100	0.8229	4.347
323.15	0.1	0.7629	1.249
323.15	20	0.7774	1.668
323.15	40	0.7896	2.099
323.15	60	0.8002	2.590
323.15	80	0.8097	3.077
323.15	100	0.8182	3.600
333.15	0.1	0.7557	1.077
333.15	20	0.7710	1.437
333.15	40	0.7837	1.789
333.15	60	0.7947	2.186
333.15	80	0.8043	2.622
333.15	100	0.8129	3.092
343.15	0.1	0.7486	0.942
343.15	20	0.7646	1.237
343.15	40	0.7779	1.527
343.15	60	0.7893	1.857
343.15	80	0.7993	2.226
343.15	100	0.8082	2.633
353.15	0.1	0.7414	0.837
353.15	20	0.7583	1.096
353.15	40	0.7721	1.353
353.15	60	0.7840	1.642
353.15	80	0.7943	1.961
353.15	100	0.8035	2.309

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.125, x_t = 0.750, x_h = 0.125$			
293.15	0.1	0.7800	1.990
293.15	20	0.7924	2.582
293.15	40	0.8032	3.245
293.15	60	0.8127	4.014
293.15	80	0.8213	4.988
293.15	100	0.8291	6.017
303.15	0.1	0.7728	1.642
303.15	20	0.7859	2.159
303.15	40	0.7971	2.685
303.15	60	0.8070	3.288
303.15	80	0.8158	3.966
303.15	100	0.8237	4.690
313.15	0.1	0.7658	1.381
313.15	20	0.7795	1.805
313.15	40	0.7912	2.227
313.15	60	0.8014	2.711
313.15	80	0.8105	3.257
313.15	100	0.8187	3.848
323.15	0.1	0.7586	1.178
323.15	20	0.7730	1.544
323.15	40	0.7852	1.902
323.15	60	0.7958	2.307
323.15	80	0.8053	2.757
323.15	100	0.8138	3.252
333.15	0.1	0.7514	1.017
333.15	20	0.7666	1.311
333.15	40	0.7794	1.617
333.15	60	0.7903	1.966
333.15	80	0.8000	2.356
333.15	100	0.8087	2.801
343.15	0.1	0.7442	0.893
343.15	20	0.7602	1.153
343.15	40	0.7735	1.415
343.15	60	0.7849	1.715
343.15	80	0.7949	2.052
343.15	100	0.8039	2.437
353.15	0.1	0.7370	0.789
353.15	20	0.7538	0.999
353.15	40	0.7677	1.221
353.15	60	0.7795	1.478
353.15	80	0.7899	1.769
353.15	100	0.7991	2.088

Table I. (Continued)

T (K)	P (MPa)	ρ (g · cm ⁻³)	η (mPa · s)
$x_m = 0.250, x_t = 0.125, x_h = 0.625$			
293.15	0.1	0.8146	2.903
293.15	20	0.8269	4.016
293.15	40	0.8376	5.412
293.15	60	0.8470	7.194
293.15	80	0.8555	9.440
293.15	100	0.8631	12.299
303.15	0.1	0.8078	2.333
303.15	20	0.8207	3.197
303.15	40	0.8318	4.254
303.15	60	0.8416	5.608
303.15	80	0.8503	7.320
303.15	100	0.8581	9.405
313.15	0.1	0.8009	1.925
313.15	20	0.8145	2.606
313.15	40	0.8260	3.423
313.15	60	0.8362	4.444
313.15	80	0.8452	5.702
313.15	100	0.8533	7.207
323.15	0.1	0.7940	1.612
323.15	20	0.8083	2.161
323.15	40	0.8203	2.821
323.15	60	0.8308	3.627
323.15	80	0.8401	4.598
323.15	100	0.8485	5.764
333.15	0.1	0.7871	1.376
333.15	20	0.8021	1.861
333.15	40	0.8146	2.388
333.15	60	0.8254	3.028
333.15	80	0.8349	3.795
333.15	100	0.8435	4.728
343.15	0.1	0.7803	1.187
343.15	20	0.7959	1.594
343.15	40	0.8089	2.042
343.15	60	0.8202	2.579
343.15	80	0.8301	3.215
343.15	100	0.8390	3.975
353.15	0.1	0.7732	1.037
353.15	20	0.7897	1.398
353.15	40	0.8033	1.787
353.15	60	0.8150	2.245
353.15	80	0.8252	2.777
353.15	100	0.8343	3.380

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.250, x_t = 0.250, x_h = 0.500$			
293.15	0.1	0.8110	2.621
293.15	20	0.8233	3.545
293.15	40	0.8340	4.728
293.15	60	0.8435	6.130
293.15	80	0.8520	7.920
293.15	100	0.8597	10.100
303.15	0.1	0.8043	2.122
303.15	20	0.8172	2.867
303.15	40	0.8283	3.730
303.15	60	0.8381	4.799
303.15	80	0.8468	6.102
303.15	100	0.8546	7.674
313.15	0.1	0.7974	1.757
313.15	20	0.8110	2.362
313.15	40	0.8225	3.082
313.15	60	0.8327	3.922
313.15	80	0.8417	4.882
313.15	100	0.8498	5.966
323.15	0.1	0.7904	1.486
323.15	20	0.8047	1.989
323.15	40	0.8168	2.535
323.15	60	0.8273	3.182
323.15	80	0.8365	3.935
323.15	100	0.8449	4.800
333.15	0.1	0.7834	1.272
333.15	20	0.7984	1.701
333.15	40	0.8110	2.161
333.15	60	0.8218	2.703
333.15	80	0.8313	3.331
333.15	100	0.8399	4.051
343.15	0.1	0.7764	1.109
343.15	20	0.7922	1.473
343.15	40	0.8052	1.859
343.15	60	0.8165	2.274
343.15	80	0.8262	2.777
343.15	100	0.8349	3.390
353.15	0.1	0.7693	0.969
353.15	20	0.7858	1.304
353.15	40	0.7995	1.622
353.15	60	0.8112	1.996
353.15	80	0.8213	2.427
353.15	100	0.8303	2.918

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.250$, $x_t = 0.375$, and $x_h = 0.375$			
293.15	0.1	0.8089	2.396
293.15	20	0.8212	3.161
293.15	40	0.8319	4.063
293.15	60	0.8414	5.159
293.15	80	0.8498	6.470
293.15	100	0.8575	8.102
303.15	0.1	0.8019	1.959
303.15	20	0.8148	2.622
303.15	40	0.8259	3.358
303.15	60	0.8358	4.239
303.15	80	0.8445	5.277
303.15	100	0.8525	6.487
313.15	0.1	0.7950	1.625
313.15	20	0.8085	2.154
313.15	40	0.8200	2.756
313.15	60	0.8302	3.471
313.15	80	0.8392	4.305
313.15	100	0.8474	5.266
323.15	0.1	0.7879	1.376
323.15	20	0.8021	1.842
323.15	40	0.8142	2.341
323.15	60	0.8247	2.915
323.15	80	0.8339	3.566
323.15	100	0.8422	4.291
333.15	0.1	0.7808	1.182
333.15	20	0.7958	1.560
333.15	40	0.8083	1.967
333.15	60	0.8192	2.433
333.15	80	0.8288	2.958
333.15	100	0.8373	3.542
343.15	0.1	0.7738	1.029
343.15	20	0.7895	1.352
343.15	40	0.8026	1.683
343.15	60	0.8139	2.063
343.15	80	0.8236	2.491
343.15	100	0.8325	2.968
353.15	0.1	0.7666	0.905
353.15	20	0.7831	1.199
353.15	40	0.7968	1.480
353.15	60	0.8085	1.809
353.15	80	0.8187	2.186
353.15	100	0.8278	2.612

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.250, x_t = 0.500, x_h = 0.250$			
293.15	0.1	0.8059	2.203
293.15	20	0.8182	2.877
293.15	40	0.8289	3.666
293.15	60	0.8384	4.610
293.15	80	0.8468	5.720
293.15	100	0.8545	7.013
303.15	0.1	0.7990	1.805
303.15	20	0.8118	2.357
303.15	40	0.8229	2.980
303.15	60	0.8328	3.710
303.15	80	0.8415	4.550
303.15	100	0.8495	5.505
313.15	0.1	0.7918	1.506
313.15	20	0.8054	1.988
313.15	40	0.8169	2.499
313.15	60	0.8270	3.075
313.15	80	0.8360	3.712
313.15	100	0.8442	4.404
323.15	0.1	0.7847	1.281
323.15	20	0.7989	1.686
323.15	40	0.8110	2.092
323.15	60	0.8216	2.591
323.15	80	0.8309	3.126
323.15	100	0.8394	3.678
333.15	0.1	0.7775	1.107
333.15	20	0.7925	1.449
333.15	40	0.8050	1.783
333.15	60	0.8159	2.167
333.15	80	0.8255	2.600
333.15	100	0.8342	3.082
343.15	0.1	0.7704	0.967
343.15	20	0.7861	1.268
343.15	40	0.7992	1.561
343.15	60	0.8104	1.897
343.15	80	0.8203	2.275
343.15	100	0.8292	2.696
353.15	0.1	0.7631	0.849
353.15	20	0.7797	1.126
353.15	40	0.7933	1.373
353.15	60	0.8051	1.656
353.15	80	0.8152	1.973
353.15	100	0.8243	2.325

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.250, x_t = 0.625, x_h = 0.125$			
293.15	0.1	0.8033	2.068
293.15	20	0.8155	2.627
293.15	40	0.8261	3.295
293.15	60	0.8356	4.083
293.15	80	0.8441	4.998
293.15	100	0.8519	6.076
303.15	0.1	0.7961	1.728
303.15	20	0.8088	2.150
303.15	40	0.8200	2.668
303.15	60	0.8298	3.267
303.15	80	0.8386	3.949
303.15	100	0.8467	4.685
313.15	0.1	0.7887	1.442
313.15	20	0.8022	1.807
313.15	40	0.8138	2.227
313.15	60	0.8239	2.713
313.15	80	0.8330	3.266
313.15	100	0.8412	3.871
323.15	0.1	0.7814	1.224
323.15	20	0.7956	1.558
323.15	40	0.8077	1.915
323.15	60	0.8183	2.323
323.15	80	0.8276	2.780
323.15	100	0.8360	3.289
333.15	0.1	0.7743	1.052
333.15	20	0.7891	1.346
333.15	40	0.8017	1.651
333.15	60	0.8126	1.998
333.15	80	0.8223	2.385
333.15	100	0.8310	2.825
343.15	0.1	0.7670	0.911
343.15	20	0.7827	1.168
343.15	40	0.7957	1.421
343.15	60	0.8071	1.725
343.15	80	0.8171	2.062
343.15	100	0.8261	2.437
353.15	0.1	0.7597	0.806
353.15	20	0.7762	1.025
353.15	40	0.7899	1.248
353.15	60	0.8016	1.492
353.15	80	0.8118	1.773
353.15	100	0.8210	2.070

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.375, x_t = 0.125, x_h = 0.500$			
293.15	0.1	0.8353	2.770
293.15	20	0.8473	3.725
293.15	40	0.8579	4.939
293.15	60	0.8672	6.481
293.15	80	0.8756	8.410
293.15	100	0.8832	10.852
303.15	0.1	0.8284	2.232
303.15	20	0.8410	2.995
303.15	40	0.8518	3.892
303.15	60	0.8616	5.019
303.15	80	0.8702	6.416
303.15	100	0.8781	8.077
313.15	0.1	0.8213	1.841
313.15	20	0.8346	2.453
313.15	40	0.8460	3.171
313.15	60	0.8560	4.053
313.15	80	0.8649	5.120
313.15	100	0.8730	6.371
323.15	0.1	0.8143	1.550
323.15	20	0.8283	2.078
323.15	40	0.8402	2.657
323.15	60	0.8506	3.347
323.15	80	0.8598	4.155
323.15	100	0.8681	5.097
333.15	0.1	0.8074	1.325
333.15	20	0.8220	1.745
333.15	40	0.8344	2.218
333.15	60	0.8451	2.771
333.15	80	0.8547	3.408
333.15	100	0.8633	4.155
343.15	0.1	0.8003	1.146
343.15	20	0.8157	1.508
343.15	40	0.8286	1.886
343.15	60	0.8397	2.335
343.15	80	0.8495	2.887
343.15	100	0.8583	3.512
353.15	0.1	0.7933	1.004
353.15	20	0.8095	1.324
353.15	40	0.8229	1.649
353.15	60	0.8344	2.038
353.15	80	0.8444	2.496
353.15	100	0.8534	3.021

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.375, x_t = 0.250, x_h = 0.375$			
293.15	0.1	0.8331	2.495
293.15	20	0.8451	3.308
293.15	40	0.8557	4.227
293.15	60	0.8650	5.446
293.15	80	0.8734	6.867
293.15	100	0.8811	8.567
303.15	0.1	0.8261	2.028
303.15	20	0.8387	2.680
303.15	40	0.8496	3.436
303.15	60	0.8593	4.336
303.15	80	0.8680	5.391
303.15	100	0.8760	6.612
313.15	0.1	0.8190	1.684
313.15	20	0.8322	2.231
313.15	40	0.8436	2.823
313.15	60	0.8537	3.529
313.15	80	0.8626	4.354
313.15	100	0.8707	5.308
323.15	0.1	0.8119	1.423
323.15	20	0.8258	1.869
323.15	40	0.8377	2.388
323.15	60	0.8480	2.952
323.15	80	0.8573	3.634
323.15	100	0.8656	4.374
333.15	0.1	0.8048	1.221
333.15	20	0.8194	1.597
333.15	40	0.8318	1.986
333.15	60	0.8426	2.440
333.15	80	0.8520	2.961
333.15	100	0.8606	3.552
343.15	0.1	0.7977	1.065
343.15	20	0.8130	1.381
343.15	40	0.8259	1.723
343.15	60	0.8371	2.122
343.15	80	0.8468	2.579
343.15	100	0.8556	3.096
353.15	0.1	0.7905	0.938
353.15	20	0.8066	1.214
353.15	40	0.8200	1.496
353.15	60	0.8316	1.820
353.15	80	0.8418	2.187
353.15	100	0.8508	2.597

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.375, x_t = 0.375, x_h = 0.250$			
293.15	0.1	0.8307	2.314
293.15	20	0.8427	3.032
293.15	40	0.8532	3.853
293.15	60	0.8626	4.838
293.15	80	0.8710	6.002
293.15	100	0.8786	7.363
303.15	0.1	0.8236	1.886
303.15	20	0.8362	2.477
303.15	40	0.8471	3.110
303.15	60	0.8569	3.866
303.15	80	0.8656	4.752
303.15	100	0.8736	5.780
313.15	0.1	0.8165	1.574
313.15	20	0.8298	2.051
313.15	40	0.8411	2.578
313.15	60	0.8512	3.192
313.15	80	0.8602	3.814
313.15	100	0.8684	4.589
323.15	0.1	0.8093	1.332
323.15	20	0.8232	1.747
323.15	40	0.8351	2.165
323.15	60	0.8456	2.648
323.15	80	0.8548	3.198
323.15	100	0.8632	3.813
333.15	0.1	0.8021	1.149
333.15	20	0.8168	1.490
333.15	40	0.8292	1.836
333.15	60	0.8399	2.238
333.15	80	0.8495	2.696
333.15	100	0.8580	3.211
343.15	0.1	0.7949	1.002
343.15	20	0.8103	1.295
343.15	40	0.8232	1.583
343.15	60	0.8344	1.916
343.15	80	0.8441	2.294
343.15	100	0.8528	2.717
353.15	0.1	0.7878	0.884
353.15	20	0.8039	1.142
353.15	40	0.8173	1.382
353.15	60	0.8289	1.659
353.15	80	0.8391	1.972
353.15	100	0.8482	2.323

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.375, x_t = 0.500, x_h = 0.125$			
293.15	0.1	0.8287	2.120
293.15	20	0.8408	2.717
293.15	40	0.8512	3.406
293.15	60	0.8606	4.217
293.15	80	0.8691	5.158
293.15	100	0.8769	6.263
303.15	0.1	0.8215	1.741
303.15	20	0.8340	2.212
303.15	40	0.8451	2.740
303.15	60	0.8548	3.348
303.15	80	0.8636	4.035
303.15	100	0.8716	4.771
313.15	0.1	0.8142	1.459
313.15	20	0.8274	1.859
313.15	40	0.8388	2.287
313.15	60	0.8489	2.768
313.15	80	0.8580	3.297
313.15	100	0.8662	3.856
323.15	0.1	0.8068	1.243
323.15	20	0.8207	1.578
323.15	40	0.8326	1.926
323.15	60	0.8431	2.325
323.15	80	0.8525	2.772
323.15	100	0.8609	3.271
333.15	0.1	0.7996	1.078
333.15	20	0.8141	1.374
333.15	40	0.8266	1.676
333.15	60	0.8374	2.024
333.15	80	0.8470	2.416
333.15	100	0.8557	2.867
343.15	0.1	0.7924	0.947
343.15	20	0.8076	1.197
343.15	40	0.8204	1.454
343.15	60	0.8318	1.746
343.15	80	0.8418	2.071
343.15	100	0.8509	2.437
353.15	0.1	0.7849	0.834
353.15	20	0.8011	1.040
353.15	40	0.8145	1.268
353.15	60	0.8262	1.517
353.15	80	0.8363	1.784
353.15	100	0.8454	2.061

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.500, x_t = 0.125, x_h = 0.375$			
293.15	0.1	0.8602	2.677
293.15	20	0.8719	3.539
293.15	40	0.8823	4.577
293.15	60	0.8915	5.890
293.15	80	0.8999	7.526
293.15	100	0.9075	9.590
303.15	0.1	0.8531	2.161
303.15	20	0.8654	2.840
303.15	40	0.8761	3.637
303.15	60	0.8857	4.618
303.15	80	0.8943	5.807
303.15	100	0.9022	7.187
313.15	0.1	0.8460	1.786
313.15	20	0.8589	2.333
313.15	40	0.8701	2.949
313.15	60	0.8800	3.709
313.15	80	0.8890	4.633
313.15	100	0.8971	5.719
323.15	0.1	0.8389	1.510
323.15	20	0.8525	1.969
323.15	40	0.8641	2.474
323.15	60	0.8744	3.077
323.15	80	0.8836	3.786
323.15	100	0.8920	4.616
333.15	0.1	0.8319	1.294
333.15	20	0.8461	1.673
333.15	40	0.8582	2.081
333.15	60	0.8688	2.561
333.15	80	0.8783	3.115
333.15	100	0.8868	3.767
343.15	0.1	0.8247	1.120
343.15	20	0.8397	1.449
343.15	40	0.8523	1.800
343.15	60	0.8633	2.211
343.15	80	0.8731	2.681
343.15	100	0.8818	3.227
353.15	0.1	0.8176	0.983
353.15	20	0.8332	1.261
353.15	40	0.8464	1.553
353.15	60	0.8577	1.890
353.15	80	0.8677	2.272
353.15	100	0.8767	2.694

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.500, x_t = 0.250, x_h = 0.250$			
293.15	0.1	0.8581	2.421
293.15	20	0.8699	3.200
293.15	40	0.8802	4.050
293.15	60	0.8895	5.081
293.15	80	0.8978	6.310
293.15	100	0.9055	7.761
303.15	0.1	0.8511	1.977
303.15	20	0.8633	2.596
303.15	40	0.8740	3.261
303.15	60	0.8837	4.047
303.15	80	0.8924	4.959
303.15	100	0.9004	6.005
313.15	0.1	0.8439	1.639
313.15	20	0.8568	2.131
313.15	40	0.8680	2.676
313.15	60	0.8780	3.303
313.15	80	0.8868	4.010
313.15	100	0.8948	4.798
323.15	0.1	0.8368	1.392
323.15	20	0.8503	1.820
323.15	40	0.8619	2.253
323.15	60	0.8723	2.748
323.15	80	0.8815	3.305
323.15	100	0.8899	3.922
333.15	0.1	0.8296	1.195
333.15	20	0.8437	1.545
333.15	40	0.8559	1.897
333.15	60	0.8666	2.298
333.15	80	0.8761	2.747
333.15	100	0.8847	3.243
343.15	0.1	0.8223	1.043
343.15	20	0.8373	1.341
343.15	40	0.8499	1.632
343.15	60	0.8609	1.967
343.15	80	0.8707	2.344
343.15	100	0.8794	2.764
353.15	0.1	0.8152	0.918
353.15	20	0.8307	1.185
353.15	40	0.8439	1.438
353.15	60	0.8554	1.725
353.15	80	0.8655	2.045
353.15	100	0.8746	2.397

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.500, x_t = 0.375, x_h = 0.125$			
293.15	0.1	0.8565	2.222
293.15	20	0.8682	2.838
293.15	40	0.8786	3.538
293.15	60	0.8879	4.382
293.15	80	0.8963	5.382
293.15	100	0.9041	6.584
303.15	0.1	0.8493	1.818
303.15	20	0.8616	2.285
303.15	40	0.8724	2.826
303.15	60	0.8821	3.456
303.15	80	0.8908	4.176
303.15	100	0.8989	4.958
313.15	0.1	0.8421	1.519
313.15	20	0.8548	1.889
313.15	40	0.8660	2.344
313.15	60	0.8760	2.864
313.15	80	0.8850	3.423
313.15	100	0.8933	3.988
323.15	0.1	0.8347	1.293
323.15	20	0.8481	1.626
323.15	40	0.8598	1.989
323.15	60	0.8702	2.398
323.15	80	0.8794	2.850
323.15	100	0.8879	3.347
333.15	0.1	0.8273	1.115
333.15	20	0.8415	1.415
333.15	40	0.8536	1.721
333.15	60	0.8643	2.068
333.15	80	0.8739	2.455
333.15	100	0.8826	2.895
343.15	0.1	0.8200	0.979
343.15	20	0.8348	1.224
343.15	40	0.8474	1.480
343.15	60	0.8585	1.771
343.15	80	0.8683	2.095
343.15	100	0.8772	2.460
353.15	0.1	0.8125	0.865
353.15	20	0.8283	1.074
353.15	40	0.8414	1.304
353.15	60	0.8528	1.559
353.15	80	0.8629	1.836
353.15	100	0.8720	2.127

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.625, x_t = 0.125, x_h = 0.250$			
293.15	0.1	0.8873	2.615
293.15	20	0.8986	3.392
293.15	40	0.9088	4.301
293.15	60	0.9178	5.438
293.15	80	0.9261	6.839
293.15	100	0.9337	8.588
303.15	0.1	0.8801	2.113
303.15	20	0.8920	2.724
303.15	40	0.9024	3.434
303.15	60	0.9119	4.285
303.15	80	0.9205	5.286
303.15	100	0.9284	6.410
313.15	0.1	0.8729	1.751
313.15	20	0.8854	2.247
313.15	40	0.8963	2.794
313.15	60	0.9061	3.449
313.15	80	0.9149	4.219
313.15	100	0.9230	5.095
323.15	0.1	0.8657	1.482
323.15	20	0.8788	1.890
323.15	40	0.8902	2.331
323.15	60	0.9003	2.848
323.15	80	0.9094	3.446
323.15	100	0.9177	4.134
333.15	0.1	0.8585	1.268
333.15	20	0.8723	1.622
333.15	40	0.8841	1.990
333.15	60	0.8945	2.417
333.15	80	0.9039	2.903
333.15	100	0.9123	3.466
343.15	0.1	0.8516	1.100
343.15	20	0.8659	1.389
343.15	40	0.8783	1.693
343.15	60	0.8891	2.042
343.15	80	0.8988	2.435
343.15	100	0.9076	2.885
353.15	0.1	0.8443	0.967
353.15	20	0.8595	1.216
353.15	40	0.8723	1.483
353.15	60	0.8835	1.775
353.15	80	0.8934	2.088
353.15	100	0.9024	2.412

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.625, x_t = 0.250, x_h = 0.125$			
293.15	0.1	0.8862	2.381
293.15	20	0.8976	3.027
293.15	40	0.9077	3.766
293.15	60	0.9168	4.638
293.15	80	0.9251	5.647
293.15	100	0.9328	6.971
303.15	0.1	0.8789	1.938
303.15	20	0.8908	2.440
303.15	40	0.9012	3.002
303.15	60	0.9107	3.668
303.15	80	0.9193	4.442
303.15	100	0.9272	5.298
313.15	0.1	0.8716	1.646
313.15	20	0.8841	2.029
313.15	40	0.8950	2.485
313.15	60	0.9047	3.016
313.15	80	0.9136	3.621
313.15	100	0.9216	4.284
323.15	0.1	0.8643	1.374
323.15	20	0.8774	1.712
323.15	40	0.8887	2.082
323.15	60	0.8989	2.512
323.15	80	0.9080	3.005
323.15	100	0.9164	3.566
333.15	0.1	0.8569	1.178
333.15	20	0.8706	1.466
333.15	40	0.8825	1.774
333.15	60	0.8930	2.131
333.15	80	0.9023	2.537
333.15	100	0.9108	3.008
343.15	0.1	0.8496	1.029
343.15	20	0.8640	1.256
343.15	40	0.8763	1.509
343.15	60	0.8872	1.799
343.15	80	0.8968	2.123
343.15	100	0.9056	2.490
353.15	0.1	0.8422	0.907
353.15	20	0.8573	1.109
353.15	40	0.8702	1.339
353.15	60	0.8814	1.594
353.15	80	0.8915	1.871
353.15	100	0.9005	2.161

Table I. (Continued)

T (K)	P (MPa)	ρ ($\text{g} \cdot \text{cm}^{-3}$)	η ($\text{mPa} \cdot \text{s}$)
$x_m = 0.750, x_t = 0.125, x_h = 0.125$			
293.15	0.1	0.9222	2.624
293.15	20	0.9332	3.321
293.15	40	0.9431	4.127
293.15	60	0.9520	5.108
293.15	80	0.9601	6.283
293.15	100	0.9677	7.774
303.15	0.1	0.9149	2.135
303.15	20	0.9264	2.663
303.15	40	0.9366	3.254
303.15	60	0.9459	3.980
303.15	80	0.9543	4.855
303.15	100	0.9621	5.899
313.15	0.1	0.9076	1.765
313.15	20	0.9196	2.192
313.15	40	0.9302	2.667
313.15	60	0.9398	3.228
313.15	80	0.9485	3.876
313.15	100	0.9564	4.618
323.15	0.1	0.9003	1.490
323.15	20	0.9128	1.864
323.15	40	0.9239	2.248
323.15	60	0.9339	2.700
323.15	80	0.9428	3.186
323.15	100	0.9511	3.768
333.15	0.1	0.8930	1.273
333.15	20	0.9061	1.569
333.15	40	0.9176	1.901
333.15	60	0.9279	2.254
333.15	80	0.9371	2.705
333.15	100	0.9456	3.200
343.15	0.1	0.8856	1.106
343.15	20	0.8994	1.383
343.15	40	0.9113	1.668
343.15	60	0.9220	1.971
343.15	80	0.9315	2.284
343.15	100	0.9402	2.629
353.15	0.1	0.8782	0.965
353.15	20	0.8926	1.220
353.15	40	0.9051	1.447
353.15	60	0.9161	1.696
353.15	80	0.9259	1.964
353.15	100	0.9349	2.248

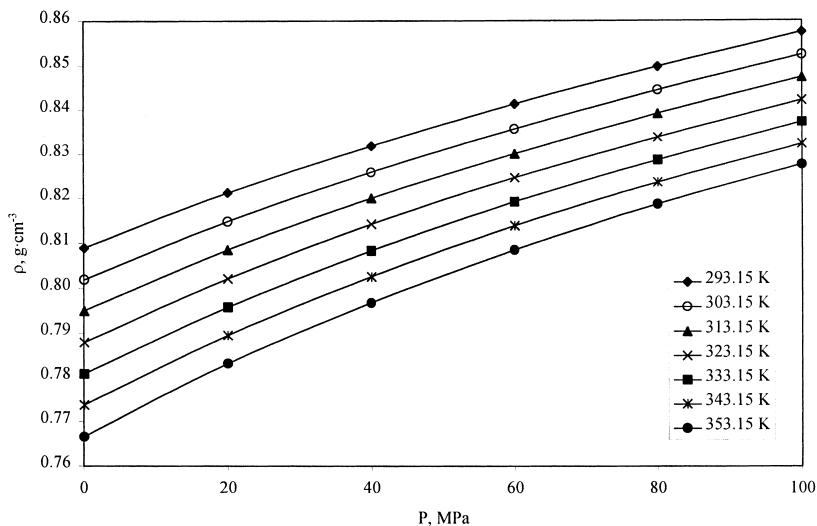


Fig. 2. Density ρ versus pressure P at various temperatures T for the composition $x_m = 0.250$, $x_t = 0.375$, $x_h = 0.375$.

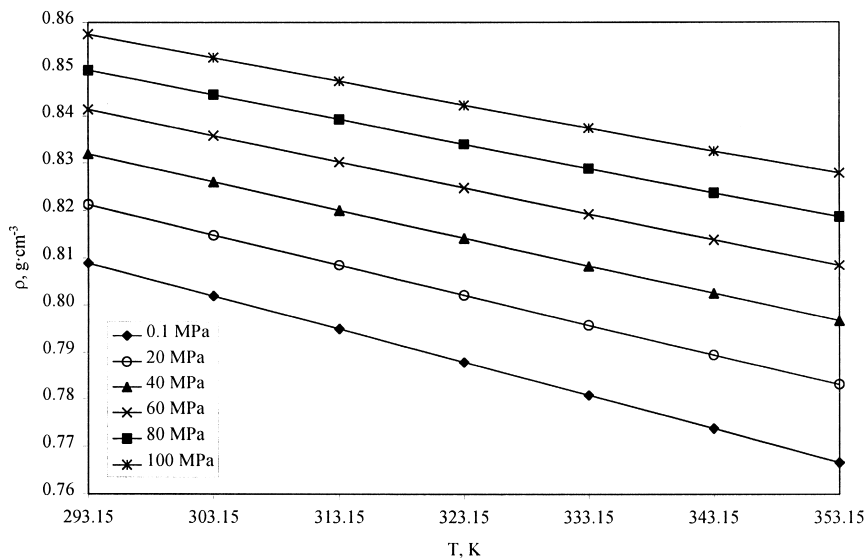


Fig. 3. Density ρ versus temperature T at various pressures P for the composition $x_m = 0.250$, $x_t = 0.375$, $x_h = 0.375$.

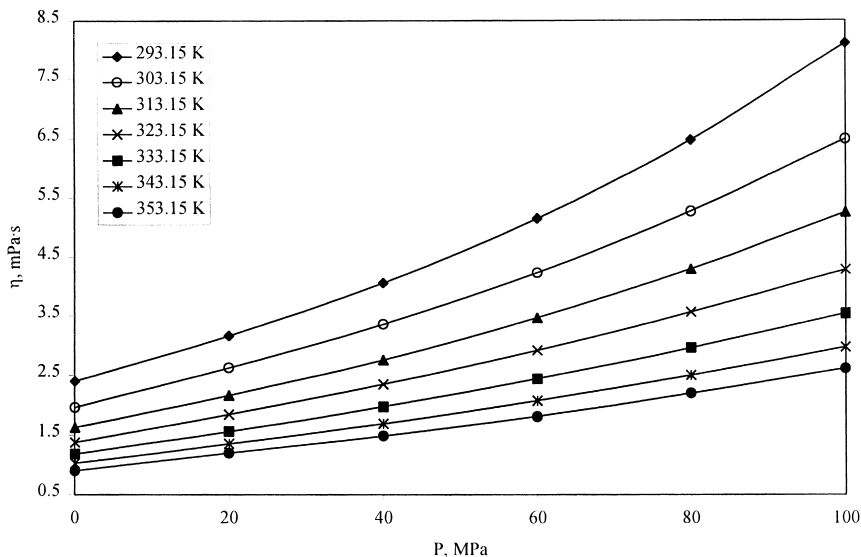


Fig. 4. Dynamic viscosity η versus pressure P at various temperatures T for the composition $x_m = 0.250$, $x_t = 0.375$, $x_h = 0.375$.

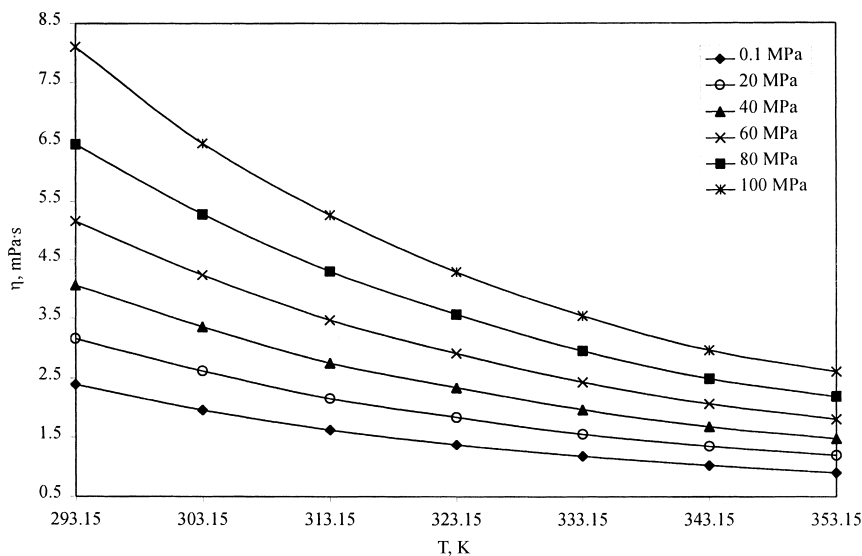


Fig. 5. Dynamic viscosity η versus temperature T at various pressures P for the composition $x_m = 0.250$, $x_t = 0.375$, $x_h = 0.375$.

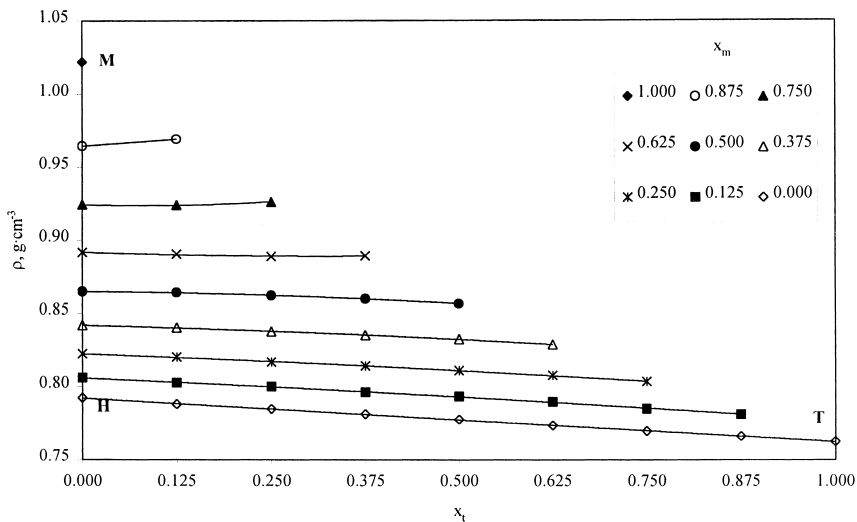


Fig. 6. Density ρ versus x_1 (n -tridecane content) for various x_m (1-methylnaphthalene content) at $P = 40$ MPa and $T = 323.15$ K.

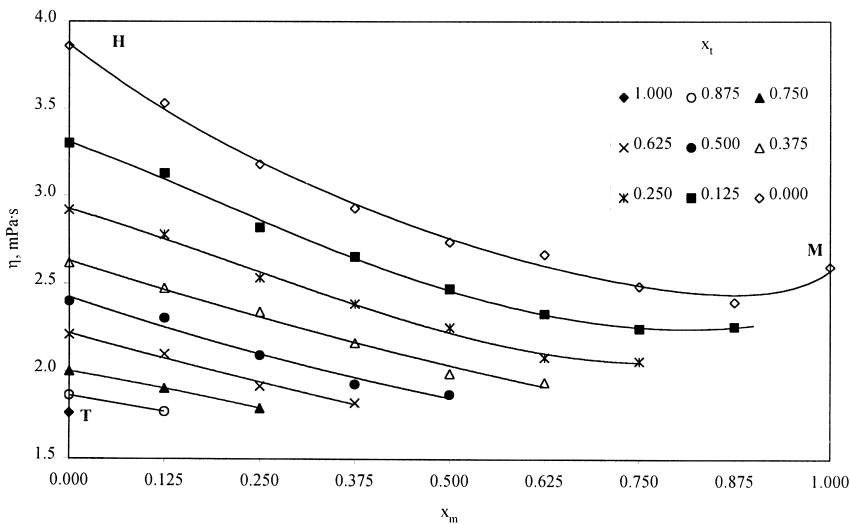


Fig. 7. Dynamic viscosity η versus x_m (1-methylnaphthalene content) for various x_1 (n -tridecane content) at $P = 40$ MPa and $T = 323.15$ K.

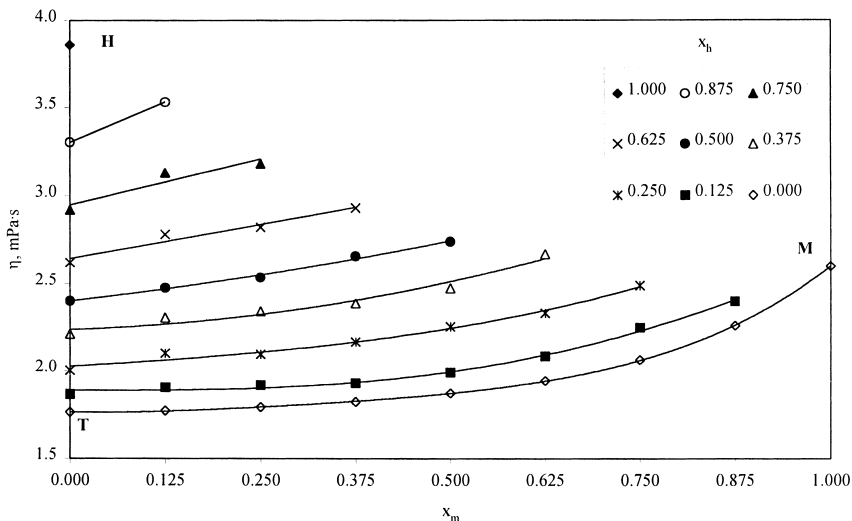


Fig. 8. Dynamic viscosity η versus x_m (1-methylnaphthalene content) for various x_h (2,2,4,4,6,8,8-heptamethylnonane content) at $P = 40$ MPa and $T = 323.15$ K.

Figure 7 shows the variations of viscosity with x_m for constant x_t at $P = 40$ MPa and $T = 323.15$ K, and Fig. 8 shows another plot as it corresponds to η as a function of x_m for constant x_h at $P = 40$ MPa and $T = 323.15$ K. Figures 9 and 10 present the viscosity surface and density surface in a ternary representation for $P = 40$ MPa and $T = 323.15$ K. It should be noted that near the 1-methylnaphthalene+2,2,4,4,6,8,8-heptamethylnonane side at a given P, T , the curves reveal a nonmonotonic behavior with respect to the composition which may be the effect of repulsive interactions (as for the binary; see Ref. 6). The minimum disappears when the amount of x_t increases. At a given composition Table I shows that it also disappears when the pressure increases. Furthermore, by keeping the concentration of 2,2,4,4,6,8,8-heptamethylnonane constant for $x_h \leq 0.375$ and plotting the viscosity as a function of the concentration of 1-methylnaphthalene for $0 < x_m < 0.500$, a very slow increase in the viscosity is observed: see Fig. 8. Overall, for the viscosity of this ternary system, there is a negative deviation from an ideal mixture, indicating that the viscosity is probably influenced by “repulsive interactions.”

4. DISCUSSION

The data obtained for the ternary system in the course of this investigation, combined with those obtained previously [4–6] on the three pure

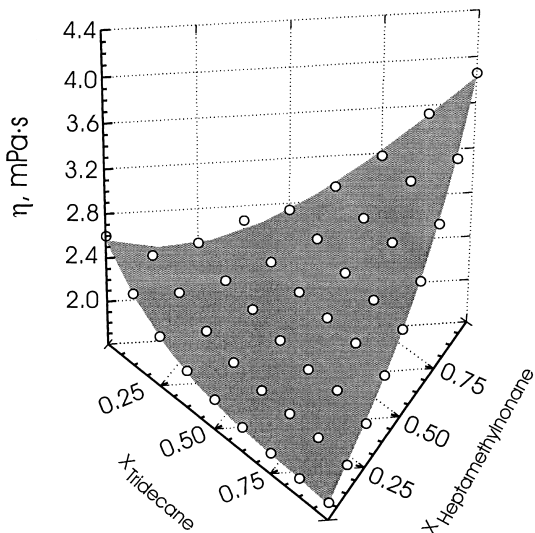


Fig. 9. Surface of the dynamic viscosity $\eta(x_m, x_t, x_h)$ in the ternary diagram at $P = 40$ MPa and $T = 323.15$ K.

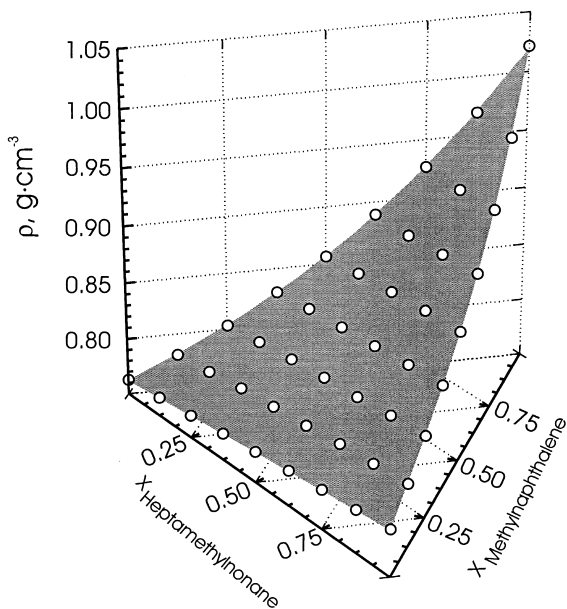


Fig. 10. Surface of the density $\rho(x_m, x_t, x_h)$ in the ternary diagram at $P = 40$ MPa and $T = 323.15$ K.

compounds and the three binaries, represent the most comprehensive set of experimental data for the dynamic viscosity (a total of 1890 data points), which can be used to test different representative models incorporating the effects of temperature, pressure, and composition. A preliminary study has been carried out focusing on mixing laws. In the case of the Grunberg and Nissan mixing law [16],

$$\ln(\eta) = x_m \ln(\eta_m) + x_t \ln(\eta_t) + x_h \ln(\eta_h) \quad (2)$$

an absolute average deviation (AAD) of 13.0% and a absolute maximum deviation (MD) of 34.2% are obtained. This model is very simple since no adjustable parameters are required, only the viscosity of the pure compounds and the composition. But with this relationship, the variation of the viscosity with composition becomes monotonic and any interaction between the pure compounds influencing the total viscosity is not taken into account. The Grunberg–Nissan mixing law can be modified by introducing adjustable parameters believed to be representative in some way of the interactions within the system studied. Using the following relationship, in which a corrective term has been added,

$$\ln(\eta) = x_m \ln(\eta_m) + x_t \ln(\eta_t) + x_h \ln(\eta_h) + (x_m x_t + x_t x_h + x_m x_h) d \quad (3)$$

each binary pair is involved in the same manner. By minimizing the least squares, the adjustable parameter ($d = -0.41484$) in Eq. (3) has been obtained, resulting in an AAD=2.92% and an MD=19.0%. Equation (2) can be made less symmetrical by introducing three parameters instead of one:

$$\ln(\eta) = x_m \ln(\eta_m) + x_t \ln(\eta_t) + x_h \ln(\eta_h) + (x_m x_t d_{mt} + x_t x_h d_{th} + x_m x_h d_{mh}) \quad (4)$$

In Eq. (4) d_{ij} is characteristic of the intermolecular interactions between component i and component j . The d_{ij} parameter can be evaluated using viscosity data for only the binary $i+j$ system. Using the three parameters $d_{mt} = -0.58110$, $d_{mh} = -0.50034$, and $d_{th} = -0.28387$ evaluated from the binary systems and reported in Ref. 17, an AAD=2.45% and a MD=15.7% are obtained for the ternary. The result is satisfactory, in the sense that the AAD is of the same order of magnitude as the experimental error. Furthermore, it can be seen from the values of the d_{ij} parameters that the largest binary interactions are obtained between 1-methylnaphthalene and n -tridecane or 2,2,4,4,6,8,8-heptamethylnonane (aromatic hydrocarbon + alkane) (see Figs. 7 and 8), rather than between n -tridecane and 2,2,4,4,6,8,8-heptamethylnonane (alkane + alkane).

Another interesting property, which can be obtained from the measured values of the viscosity and density, is the excess activation energy of viscous flow ΔG^E , which appears in

$$\ln(\eta V) = x_t \ln(\eta_t V_t) + x_m \ln(\eta_m V_m) + x_h \ln(\eta_h V_h) + \Delta G^E / RT \quad (5)$$

where R is the gas constant and $V_i = M_i / \rho_i$ is the molar volume of component i . M_i is the molecular weight of component i , and for a mixture the molecular weight $M = \sum x_i M_i$. This relation is a modified form of the equation of Katti and Chaudhri [18] and is theoretically justified by Eyring's representation of the dynamic viscosity of a pure fluid [19]. It is important to note here that the quantity ηV is also obtained from the time-correlation expression for shear viscosity [20]. Thus, the quantities ηV and ΔG^E have a theoretical background, while the corrective terms in Eqs. (3) and (4) do not. The excess activation energy of viscous flow ΔG^E can be calculated from the results in Table I and in Refs. 4–6. Figure 11 shows the surface $\Delta G^E(x_m, x_t, x_h)$ in a ternary representation for $P = 40$ MPa and $T = 323.15$ K. We note that ΔG^E is generally negative and $|\Delta G^E|$ increases with pressure. For some authors [21, 22] the fact that the excess activation energy of viscous flow ΔG^E is negative means that the predominant effect in the mixture is the breaking-up of the ordered structure present in the pure liquids. Other authors [23, 24] interpret the negative values of ΔG^E by the fact that the repulsive forces of interaction are the forces which predominate, corresponding to the breaking of bonds within the ordered structure. For the very associative system water + alcohol (see, e.g., Ref. 11) $|\Delta G^E|$ can reach $5000 \text{ J} \cdot \text{mol}^{-1}$, whereas in this work the maximum value of $|\Delta G^E|$ is about $600 \text{ J} \cdot \text{mol}^{-1}$, which corresponds to weak interactions and consequently to a weakly interactive system. The estimation of the excess activation energy of viscous flow is based on Eq. (5), which can be thought to model the viscosity of an ideal mixture by assuming $\Delta G^E = 0$ and, therefore, using only the properties of the pure components. In this case an AAD=8.65% and a MD=27.7% are obtained. This relationship can also be used to model the viscosity of real mixtures. We tried $\Delta G^E = (x_m x_t + x_m x_h + x_t x_h) W$ in which W is an adjustable parameter representative of the interactions. By least-squares minimization, we obtained $W = -788.13 \text{ J} \cdot \text{mol}^{-1}$ resulting in an AAD=2.12% and a MD=16.8%. We also tried $\Delta G^E = (x_m x_t W_{mt} + x_m x_h W_{mh} + x_t x_h W_{th})$, where W_{ij} is characteristic of intermolecular interactions between component i and component j responsible for the excess energy of activation for viscous flow. The W_{ij} parameters can be evaluated using viscosity and density data concerning only the binary $i+j$ system. For each of the three binary $i+j$ systems we found $W_{mt} = -1173.6 \text{ J} \cdot \text{mol}^{-1}$, $W_{mh} = -611.55 \text{ J} \cdot \text{mol}^{-1}$, and $W_{th} = -726.83 \text{ J} \cdot \text{mol}^{-1}$ by

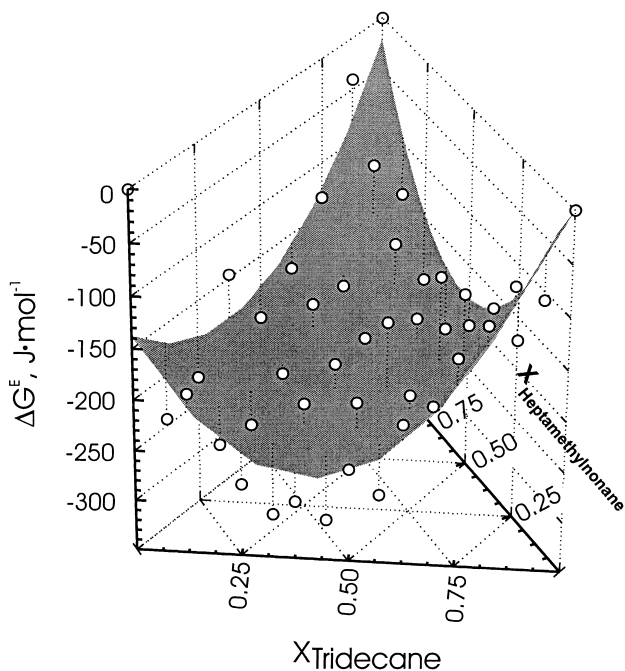


Fig. 11. Surface of excess activation energy of viscous flow $\Delta G^E(x_m, x_i, x_h)$ in the ternary diagram at $P = 40$ MPa and $T = 323.15$ K.

least squares. Using these three W_{ij} parameters results in an AAD = 2.12% and a MD = 14.6% for the ternary mixtures. The W_{ij} parameters are related to the excess activation energy of viscous flow and they are all negative, indicating that repulsive forces dominate between component i and component j .

5. CONCLUSION

The viscosity and density measurements reported in this work together with the viscosity and density measurements for the pure components and binary mixtures [4–6] represent the most comprehensive study for the ternary system 1-methylnaphthalene + n -tridecane + 2,2,4,4,6,8,8-heptamethylnonane. This study is a part of a more general study concerning various systems (associative and nonassociative mixtures, various binary mixtures with different compounds, ternary systems, and even systems with more than three components). We intend to undertake a modeling study of these measurements concerning three recently proposed viscosity models:

the hard-sphere viscosity scheme [25, 26], a free-volume viscosity model [27–29], and the viscosity model based on the friction theory [30, 31]. Each of these three models has a strong physical background and they are conceptually very different.

ACKNOWLEDGMENT

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